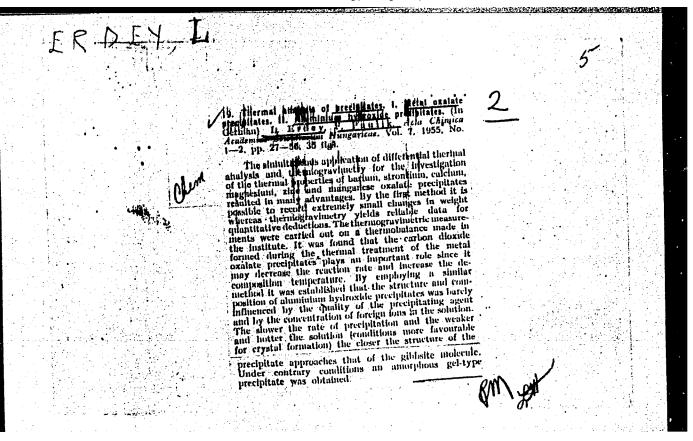
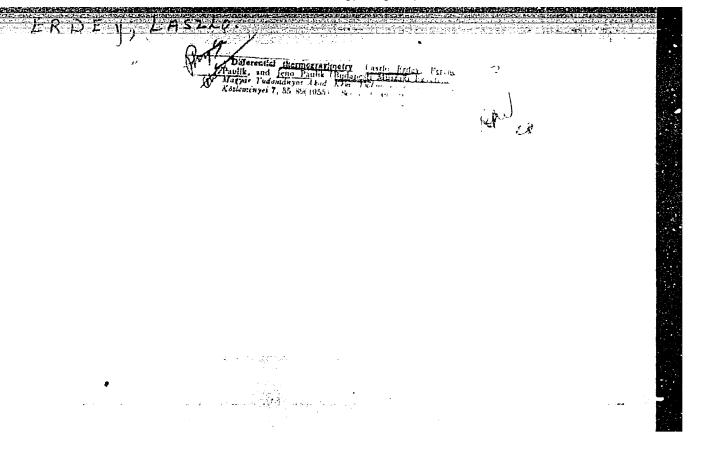
ERDEY, L.

Report of the work of the Chemical Section; also, remarks by G. Schay and others. p. 3. KOZLEMENYEI. Budapest. Vol. 7, no. 1, 1955.

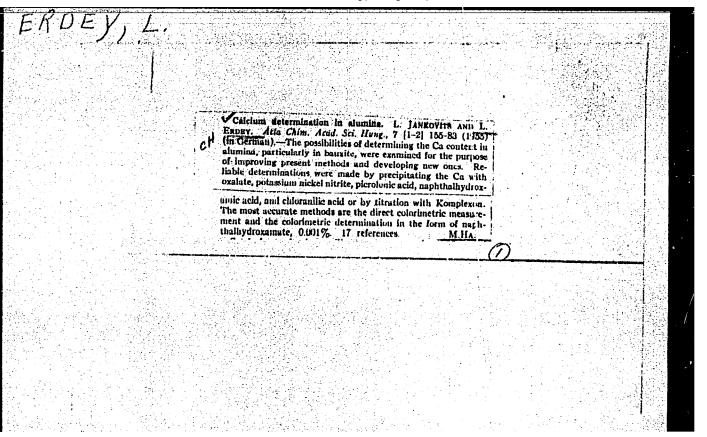
SOURCE: East European Accessions List (FEAL), LC, Vol. 5, No. 2, Feb. 1956

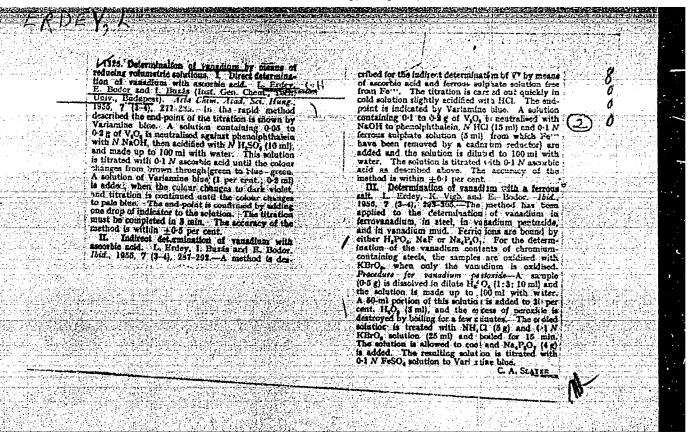




21. Data on the kinetics of the decomposition of hydrogen
peroxide in alkaline medium, (Ito German) I. Wricy
1. Incaédy. Acto. Chimica Academila Stentiarum
Himitaritate Vol. 7. 1955. No. 1-2. pp. 93-115. Thes.
3 talas.

Inased on a theoretically derived reaction equation
the decomposition process proved to be of the second
orother. Reperimentally however, the second order was
found only at the pit yalne of maximum decomposition.
The activation energy of the decomposition was calculated from the rate constants established at different
temperatures for the pit yalnes of maximum decomposition. The activation energy proved to be independent
of the dimensions of the endosing glass surface. A linear
relationality was found to exist between the decomposition
rate and the surface area celow at maction of the square
the decomposition. The alkaline decomposition of
laydrogen peroxide was influted by a starting period
possibly due to the formation of this intermediate.
The starting period
possibly due to the formation of a hypothetical Intermediate. It is the deformation of this intermediate
well known decomposition products.





ERDEY, L.; VIGH, K.; BODOR, E.

ERDEY, L.; VICH, K.; BODOR, E. Determination of vanadium using reducing measuring solutions. III. Determination of vanadium with an iron (II) salt as a measuring solution. In German. p. 293.

Vol. 7, no. 3/4, 1955 ACTA CHIMICA SCIENCE HUNGARY

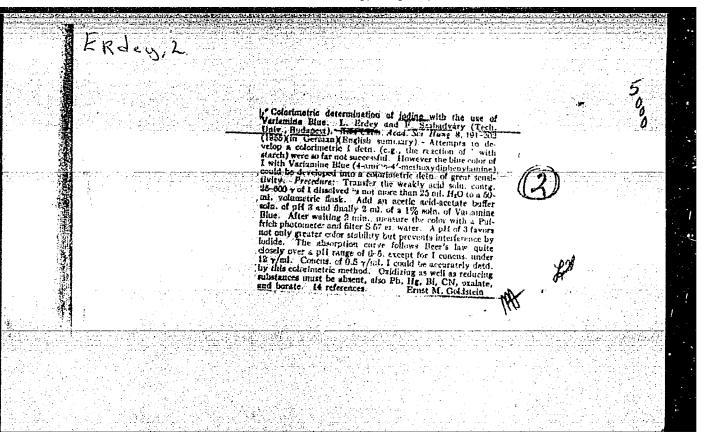
So: East Europeon Accessions, Vol. 5, No. 9, Sept. 1956

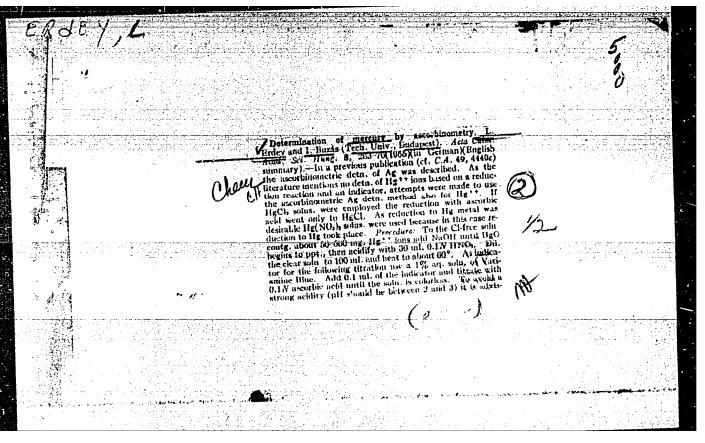
ERDEY, L.; GEGUS, E.; KOCSIS, E.

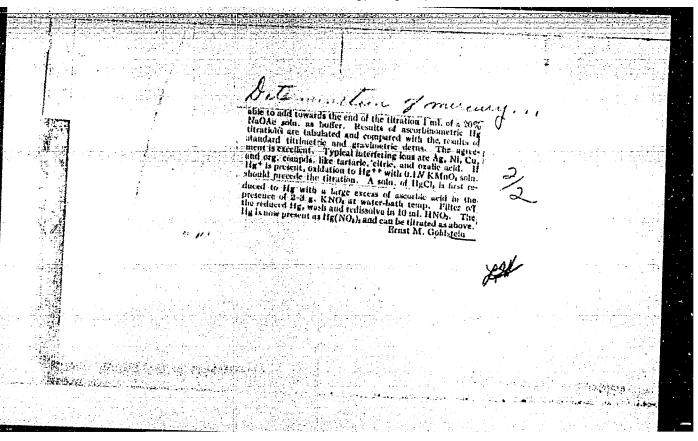
ERDEY, L.; GEGUS, E.; KOCSIS, E. Spectral analysis of solutions using the cup electrode method. In German. p. 343.

Vol. 7, no. 3/4, 1955 ACTA CHIMICA SCIENCE HUNGARY

So: East Europeon Accessions, Vol. 5, No. 9, Sept. 1956

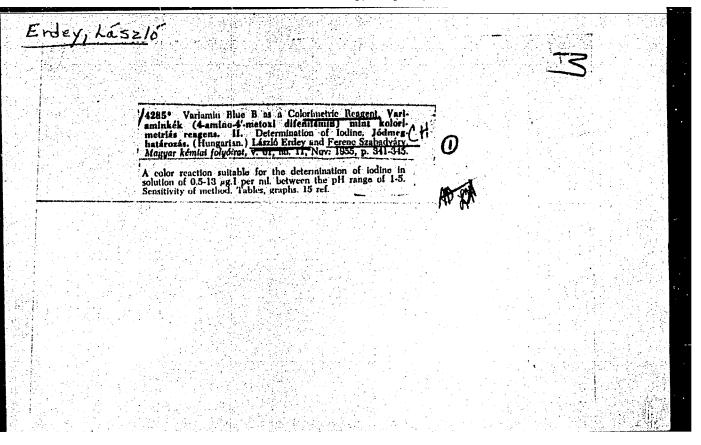






"APPROVED FOR RELEASE: Thursday, July 27, 2000

CIA-RDP86-00513R00041221



EMPEY, L.; FUZAS, L.

ENDEY, L.; BUZAS, L. Easily produced analytic glass filters. p. 443.

Vol. 611 No. 12, Dec. 1955. MACYAR KEMIAI FOLYCIRAT SCIENCE Eudapest, Hungary

So: East European Accession, Vol. 5, No. 5, May 1956

ERDEY, Lasylo

HUNGARY/ Analytical Chemistry. General Problems. G-1

Abs Jour: Referat. Zhur.-Khimiya, No. 8, 1957, 27137 K.

Author : Laszlo Erdey.

Title : Introduction into Chemical Analysis. Part I.

Qualitative Analysis. Textbook for Universities.

4th Edition.

Orig Pub: Budapest, Tankonyvkiado, 1956, VI, 281 1.,

28.60 ft.

Abstract: no abstract.

Card 1/1

CIA-RDP86-00513R00041221 "APPROVED FOR RELEASE: Thursday, July 27, 2000

HUNGARY/Analytical Chemistry - Analysis of Inorganic Substances.

E.

Abs Jour

: Ref Zhur - Khimiya, No 9, 1958, 28486

Author

Inst

Erdey, L. and Banyai, E.

Title

: The Utilization of Exchange Precipitation Reactions in Analytical Chemistry. II. The Determination of the

Chloride Ion.

Orig Pub

: Magyar tud akad Mem tud oszt koczl, 7, No 2, 175-186

(1956) (in Hungarian)

Abstract

: See RZhKhin, 1957, 1257.

Card 1/1

51

CIA-RDP86-00513R00041221 "APPROVED FOR RELEASE: Thursday, July 27, 2000

HUNGARY/Analytical Chemistry - Analysis of Inorganic Substances.

E.

Abs Jour

: Ref Zhur - Khimiya, No 9, 1958, 28482

Author

: Laszlo, E. and Banyai, E.

Inst

Title

: The Utilization of Exchange Precipitation Reactions in Analytical Chemistry. III. The Determination of Sulfate

and Sulfide Ions.

Orig Pub

: Magyar tud akad Kem tud oszt koezl, 7, No 2, 187-198

(1956) (in Hungarian)

Abstract

: See RZhKhim, 1957, 8534.

Card 1/1

HUNGARY/Analysis of Inorganic Substances.

G-2

Abs Jour: Ref Zhur-Khimiya, No 6, 1957, 19583

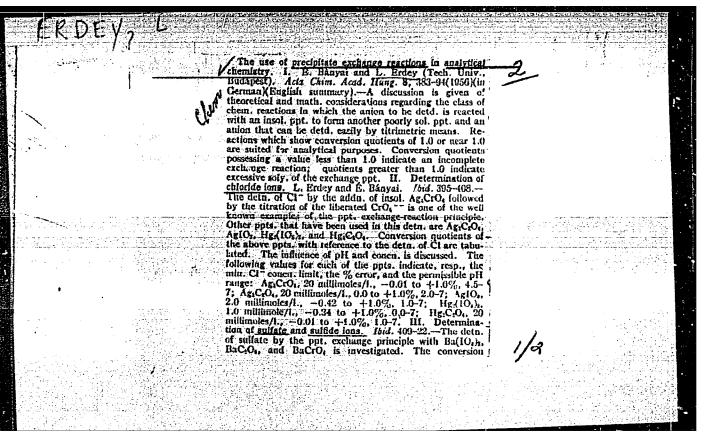
TO.104 ml. Cu, Bi and Co interfere in the molar ratio 1:1, Cd, Al and Zn interfere. F and PO₄ retard the basic reaction, NO₃ (at 1:10 NaNO₃) does not impede. The possibility of titrating the solution of complexon III with a solution of Fe³ in the presence of variamine blue in the region of pH 3 - 4.5 at 50° was established. The error is 1 - 2%.

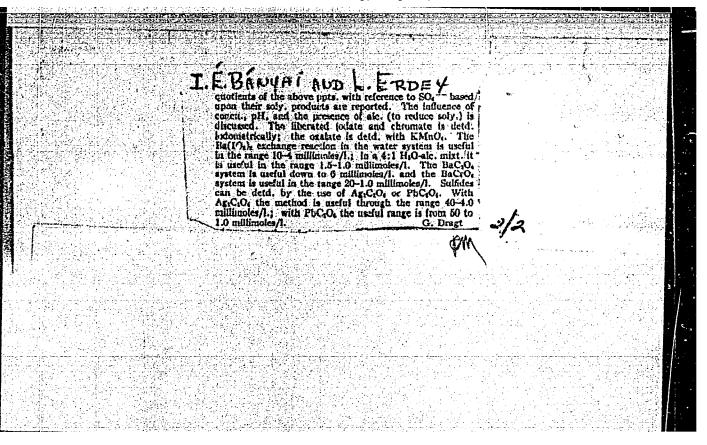
Card 2/2

- 59 -

"APPROVED FOR RELEASE: Thursday, July 27, 2000

CIA-RDP86-00513R00041221





G-2

Hungary/Analytical Chemistry - Analysis of Inorganic Substances

Abs Jour : Referat Zhur - Khimiya, No 3, 1957, 8534

view of the very low solubility of III in neutral media, the reaction with III is carried out in a 0.1 N HCl solution with refluxing, NH4OH is added until a faintly yellow color appears, and the ${\rm CrO_{l_1}}^{2-}$ determined after the separation of the precipitate from the filtrate. In agreement with theoretical culculations (RZhKhim, 1956, 78375), it has been established that the reaction with I can be applied to the determination of SOl_1^{2-} only in the concentration range 4-10 mmol/liter; the error is less than 1%. When the solubility of I is lowered by the addition of alcohol, the range of application of the reaction is shifted to the 1-1.5 mmol/ liter region. The reaction with I results in a 12-fold increase in the titer of the solution and is therefore suited for the determination of very small amounts of $SO_{l_1}^{2-}$ in neutral, weakly acidic, or ammoniacal solutions in the narrow concentration range indicated. The reaction with II can be applied to the determination of SO_{i}^{2} in neutral or ammoniacal solutions at concentrations ≥ 6 mmol/liter. The range of applicability of the reaction with III is from 1-20 mmol/liter SO_4^2 . I, II, and III are prepared by the reaction of BaCl₂ with KIO₃, (NH₄)₂C₂O₄, and K₂CrO₄. The

Card 2/3

ERDEY, L.

The development of polarography in Hungary. In German. p. 17. (Acta Chimica, Vol. 0, No. 1/4, 1956, Budapest, Hungary)

SO: Monthly List of East European Accessions (EEAL) IC, Vol. 6, No. 8, Aug 1957. Uncl.

HUNGARY/Analytical Chemistry. Analysis of Inorganic Substances.

E-2

Abs Jour: Ref Zhur-Khim., No 13, 1958, 43045.

Author : Erdey L. Karsay A.

: Hungarian Academy of Sciences. Inst

: Amperometric Determination of Ions of Trivalent Title

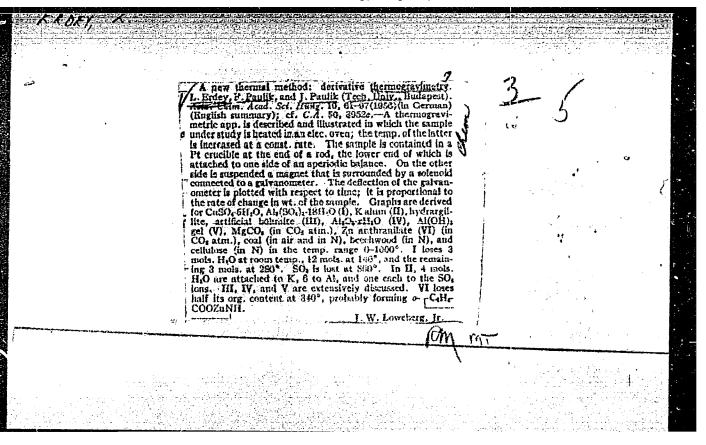
Iron With Ascorbic Acid.

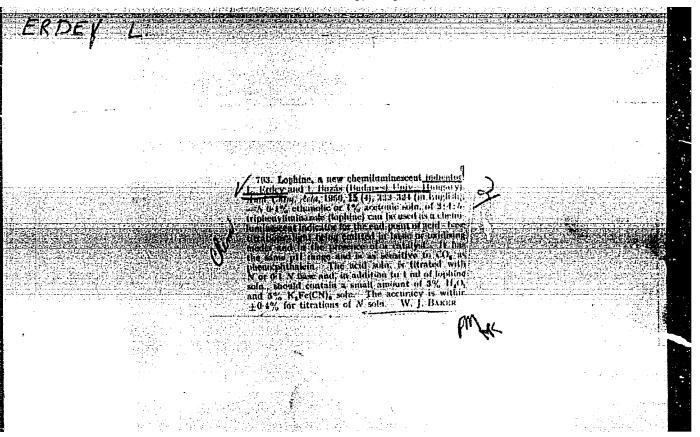
Orig Pub: Acta chim. Acad. sci. hung., 1956, 9, No 1-4, 43-48.

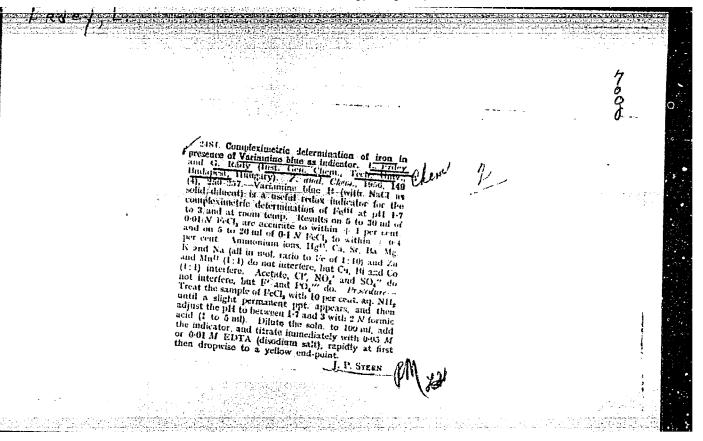
Abstract: It was found that aqueous solutions of ascorbic acid (I) can be used in amperometric titration of Fe3+at concentrations as low as 0.001 M. On determination of 1-2 mg Fe the error is less than 1% which is commesurable with the accuracy of the other known methods. The advantages of \underline{I} in comparison with other titration reagents are the ready preparation of a solution of I

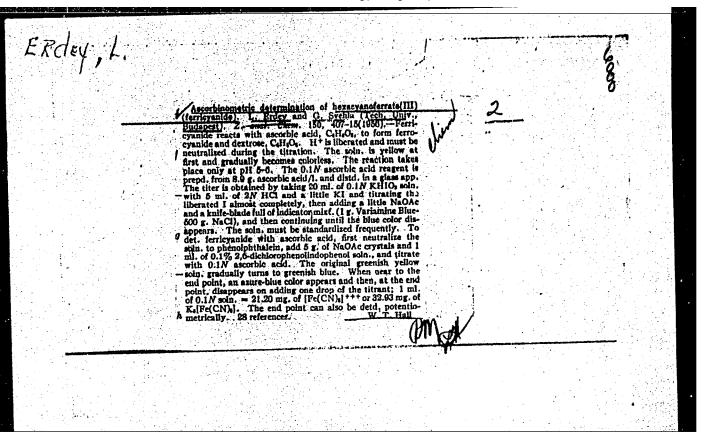
: 1/2 Card

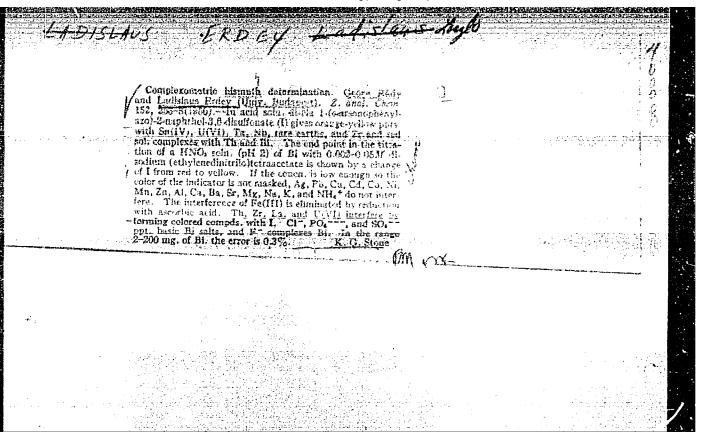
9

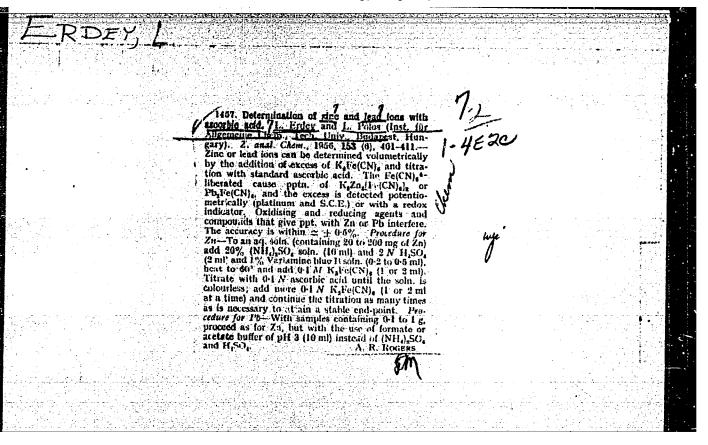


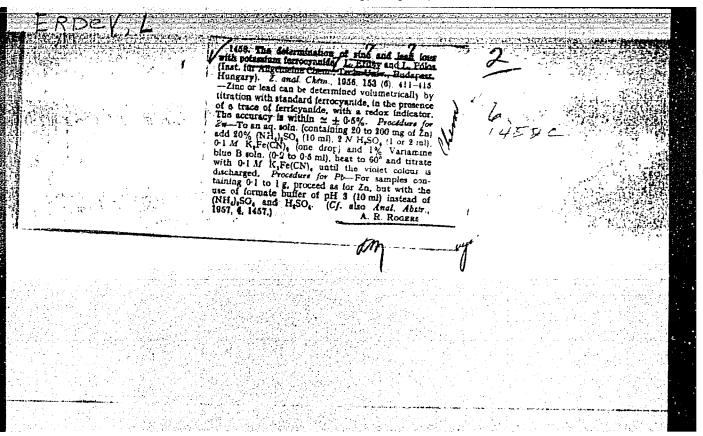


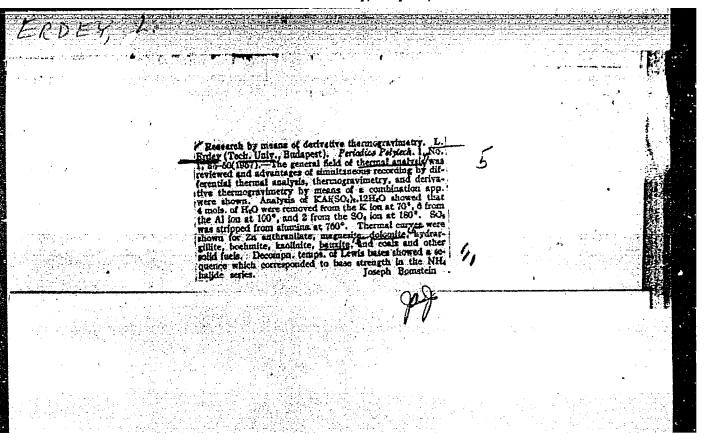


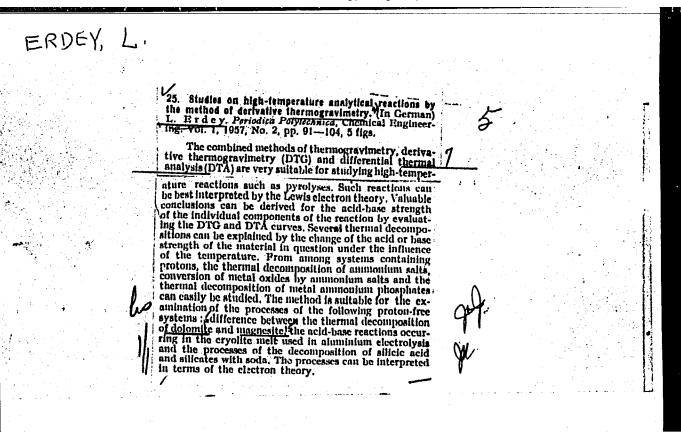












ERDEY, I.

Report on the work of the Section of Chemical Sciences; also, remarks by M. Freund and others.

p. 169 (Kozlemenyel.) Budapest Vol. 8, no. 2/3 1957

SO: Monthly Index of East European Acessions (AEEI) Vol. 6, no. 11 November 1957

ERDEY, L.; BANYAI, E.; PAULIK, E.

"The use of precipitate exchange reactions in analytical chemistry." IV.

p. 103 (Kozlemenyel) Vol. 9, no. 1, 1957 Budapest, Hungary

SO: Monthly Index of East European Accessions (EEAI) LC. Vol. 7, no. 4 April 1958

ERDEY, L,

HUNGARY/Analytical Chemistry - Analysis of Inorganic Substances. E-2

Abs Jour: Referat Zhur-KHimiya, No 5, 1958, 14174.

Author : Erdey L., Vigh K.

: Hungarian Academy of Sciences

: Permanganatometric Determination of Vanadium in Ferrovana-Inst Title

dium After Reduction with Sodium Mitrite.

Orig Pub: Acta chim. Acad. sci. hung., 1957, 11, No 1-2, 73-83;

Magyar tud. akad. Kem. tud. oszt. kozl., 1956, 7, No 2,

277-285

Abstract: To the sample of ferrovanadium are added 50 ml H2SO4 (1:1)

and 20 ml HNO3 (1:3), evaporation is carried out until SO3 vapors are formed, diluted with water to 200 ml, SiO2 is separated and solution cooled to room temperature. Decomposition of ferrovanadium can also be effected by successive treatment with 50 ml H2SO4 (1:1) and 5-10 ml 30 H2O2. To

: 1/2 Card

HUNGARY/Analytical Chemistry - Analysis of Inorganic Substances.

E-2

Abs Jour: Referat Zhur-Khimiya, No 5, 1958, 14174.

the resulting solution is added 1 g NaNO₂, stirred, after 10 minutes 1.5 g of urea are added, heated to 60-70° and titrated with 0.1 N solution of KMnO₁. A control experiment is run concurrently. Satisfactory results were obtained.

Card : 2/2

ERDEY 1

HUNGARY/Analytical Chemistry - Analysis of Inorganic Substances. E-2

Abs Jour: Referat Zhur-Khimiya, No 5, 1958, 14155.

Author : Erdey L., Karsai A.

: Hungarian Academy of Sciences Tnst

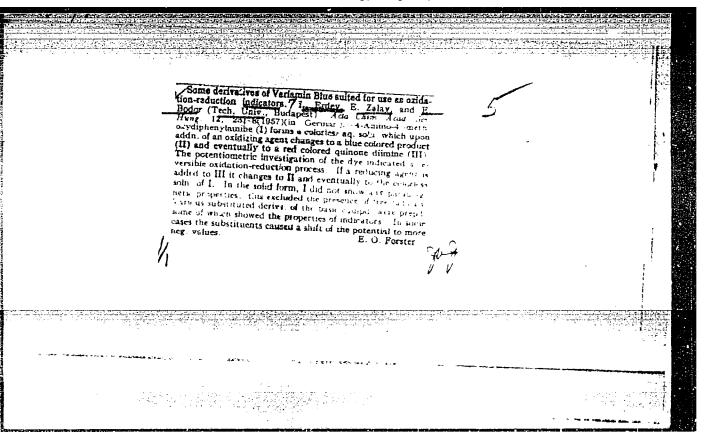
: Indirect Method of Polarographic Determination of Calcium. Title

Orig Pub: Acta chim. Acad. sci. hu g., 1957, 11, No 1-2, 171-178.

Abstract: Description of a method for determining 6.3 . 10-4 to 2 . 10-2 mole/liter Ca, which is based on precipitation

of Ca with bromanilic acid (I) and a subsequent determination of excess I, which is reduced polarographically at pH 4.5 and has an $E_{1/2} = 0.21$ v (in relation to a saturated calomel electrode). On carrying out the analysis 5 ml 0.1% solution of I are mixed with 0.5-4 ml of a solution of Ca and after 10 mimutes are added 5 ml 1 M CH3COOH containing 3 ml 2 M NH4Cl in 50 ml solution; No is passed for 5 minutes and polarography is carried out. Under the same condition the polarogram of

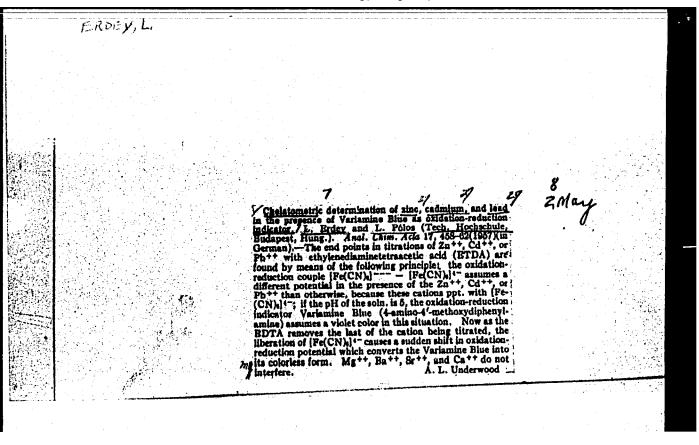
: 1/2 Card

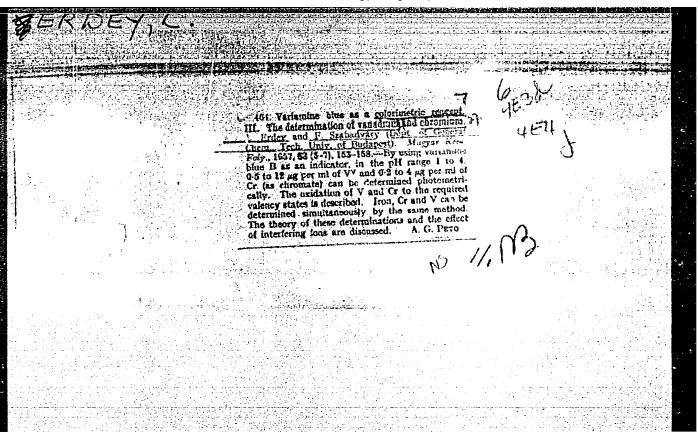


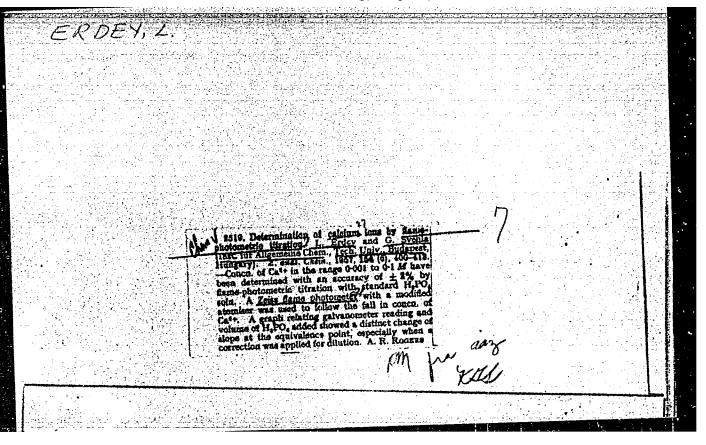
ERDEY, L.

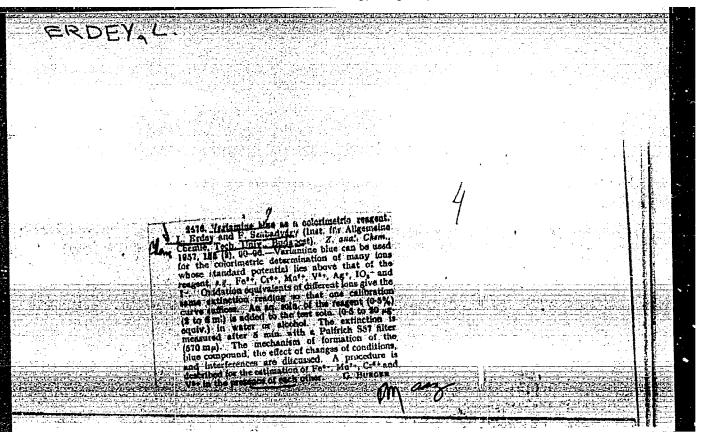
16. Recent results of derivative thermogravimetry. (In German) P. Paullk, L. Brdey. Acta Chimica Academiae Scientiarum Hungaricus. Vol. 13, 1957, No. 1-2, pp. 117-140, 19 figs.

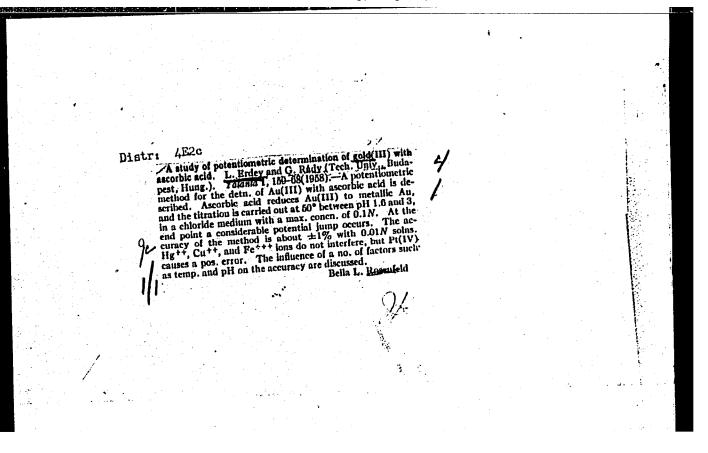
Investigations carried out so far by the method of derivative thermogravimetry proved that the derived curve facilitates the evaluation of the difficulty interpretable basic curves. Processes taking place in rapid sequences in the substance investigated or reactions causing very small losses of weight may readily be detected by this method with a high degree of sensitivity. The temperature of the maximum of the derived curve clearly defines the reaction under examination. If the values of the initial and final temperatures of the reaction are plotted on the basic curve precise stoichiometric calculation can be made. Very significant conclusions can be drawn from the comparison of the curves obtained by derivative thermogravimetry with those by differential thermogravimetry of various analytical precipitates, hauxites alumina hydrates, red muds, ervolites and catalysts are discussed.











COUNTRY: Poland L-1

CATEGORY

ABS. JOUR. : AZKhim., No. 1959, No. 85967

AUTHOR : Erdey, L.

INST. : Titration with the Use of Chemilumine scent

Indicators.

ORIG. FUB.: Chem. analit., 1958, 3, No 3-4, 269-280

ABSTRACT: Chemiluminescent indicators (CI) are considered as redox systems; in the process of oxidation of CI the electrons which are in excited state, emit on transition to stable state, a portion of the energy in the form of a light quantum. In presence of CI, in alkaline solutions, light quantum. In presence of CI, in alkaline solutions, spontaneous decomposition of H₂O₂ takes place. Beginning of spontaneous decomposition of H₂O₂ takes place. Beginning of reaction of CI which is associated with chemiluminescence, reaction of CI which is associated with chemiluminescence, occurs on reaching a definite pH value of the solution, or occurs on reaching a definite pH value of the solution, or occurs on reaching a definite pH value of the solution, or occurs on reaching a definite pH value of the solution, or occurs on reaching a definite pH value of the solution, or occurs on reaching a definite pH value of the solution, or occurs on reaching a definite pH value of the solution. Therefore, CI are suitable for determining end-point of acid-base as well as oxidation-reduction titrations. The advantage of CI over indicators of other types is the

CARD: 1/3

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COUNTRY
CATEGORY

ABS. JOUR. : RZKhim., No. 1959, No. 855(7)

AUTHOR
INST. :
TITLE

ORIG. PUB. :

ASSTRACT : fact that they make possible titration of turbid and colored solutions. As CI were studied lucigenin (dimethylliac idyl nitrate) (I), luminol (hydrazide of 3-(dimethylliac idyl nitrate) (I), luminol (hydrazide of 3-(dimethylliac idyl nitrate) (I), saminophthalic acid), lophin (2,4,5-triphenylimidazole), and aminophthalic acid), lophin (2,4,5-triphenylimidazole), and sminophthalic acid), lophin (2,4,5-triphenylimidazole), and sminophthalic acid), solution of strong and weak acids, and also sultable for titration of strong and weak acids, and also sultable for titration of strong and weak acids, and also sultable for titration of strong and weak acids, and also sultable for titration of strong and weak acids, and also sultable for titration of strong and weak acids, with the of strong oases. In titration of acids or bases with the of strong oases. In titration of acids or bases with the of strong oases. In titration of acids or bases with the of strong oases. In titration being titrated, acid, until the green glow of the solution vanishes. With acid, until the green glow of the solution vanishes. With acid, until the green glow of the solution vanishes. With acid, until the green glow of the solution wanishes. With acid, until the green glow of the solution vanishes.

COUNTRY : Poland E-1

CATEGORY :

ABS. JOUR. : AZKhim., Ko. 1959, Ro. 85967

AUTHOR : INST. : TITLE :

ORIG. PUB. :

ABSTRACT: of H₂O₂, and of oxidizing agents -- with solutions of N₂H₄.H₂SO₄, have also been developed. Luminol has no resersible indicator properties and can be used only for titration of acids with alkalies, in the presence of H₂O₂ and of catalysts, and also for titration of reducing agents with solutions of NaClO or NaBrO. Lophin, analogous to luminol in mechanism of luminescence, is suitable for titration of strong and weak acids. Silexen shows a red glow in the presence of strong oxidizing agents, and is recommended as CI in cerimetric, chromatometric, and permanganatometric titrations. A variant of automatic titration with the use of CI has been developed. -- A. Nemodruk.

CARD: 3/3

В

CZECHOSLOVAKIA/Physical Chemistry. Thermodynamics. Thermo-

chemistry. Equilibria. Phase Transitions. Physical-

Chemical Analysis.

Abs Jour: Ref Zhur-Khim., No 5, 1959, 14554.

Author : Erdey L.

: An Application of the Differential Thermogravimetric Inst Title

Method.

Crig Pub: Chem. zvesti, 1958, 12, No 6, 352-365.

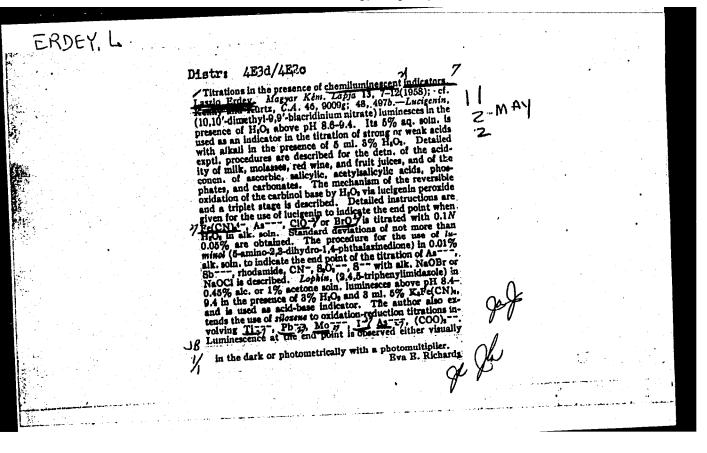
Abstract: Review and comparison of methods for differential

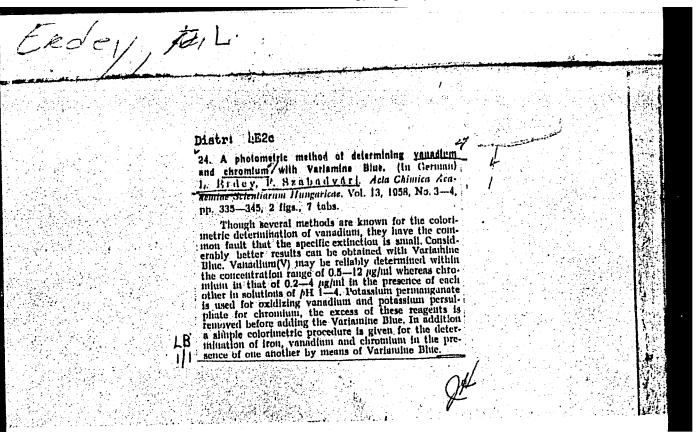
thermic analysis, thermogravimetry and differential thermogravimetry. See also Ref Zhur-Khim., 1958,

52927-52930; 57132.

: 1/1 Card

13





Distr: LE2c(j)

23. Precipitate exchange reactions in analytical chemistry, 17°. (In German) L. Erdey, B. Banyai, P. Paulik, Acta Chimica deademina Scientiarum Hungaricae. Vol. 13, 1958, No. 3-4, pp. 453-403, 8 tabs.

Bubsequent to the theoretic discussion of the exchange of chloride by mercury(ii) iodate the practical conditions of the method of determination on this basis are discussed. Between certain limits of concentration the main reaction between mercury iodate and chloride ions proceeds without any side reactions. However in solutions of higher concentration a HgCl-7 complex whereas in solutions of lower concentration a HgCl-1 complex whereas in solutions of lower concentration a HgCl-1 complex whereas in solutions of lower concentration and HgCl-1 complex in turn, more lodate than expected on the HgCl-1 complex, in turn, more lodate than expected on the basis of the main reaction. The determination of chloride may be carried out also on a micro scale in the presence; of alcohol and under adequate conditions.

Promide, iodide and eyantide ions may be similarly determined in this way.

: HUNGARY :Analytical Chemistry. General Problems Country Category No. 15039 : Ref Zhur - Khim., No 5, 1959, Abs. Jour : Erdey, L.; Banyai, E.; Zalay, R.; Tesy, M. : Hungarian Academy of Sciences Author : Preparation of Derivatives of Variamine Blue Institut. and Their Standard Oxidation-Reduction Poten-Title : Acta chim. Acad. scient. hung., 1958, 15, Orig Pub. No 1, 65-79 : A description is given of the preparation of the following derivatives of variamine blue Abstract (I) which differ from I itself according to the value of the standard oxidation-reduction potential (SORP), and which can be used as oxidation-reduction indicators (ORI) as follows: 4-amino-2-methyl-4'-methoxy-diphenylamine (II), 4-amino-4'-methoxy-diphenylamine-2-sulfo-acid (III), anilide of 4-amino-41methoxy-diphenylamine-2-sulfo-acid (IV), 1/6 card:

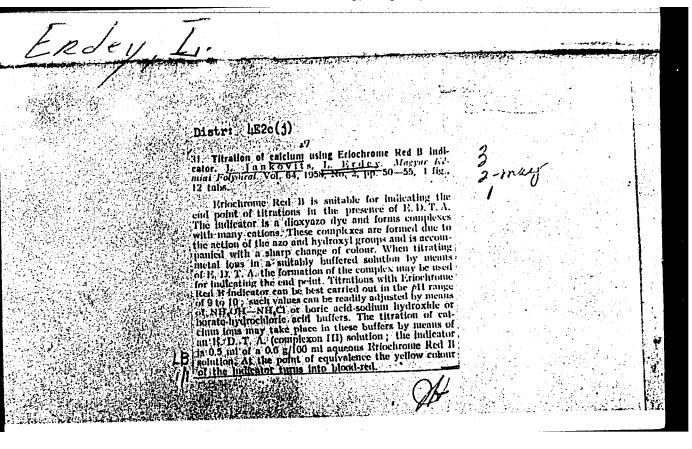
: Analytical Chemistry. General Problems Catogory No. 15039 : Ref Zhur - Khim., No 5, 1959, Abs. Jour Author Institut. Titlo Oris Pub. : anisidid of 4-amino-4'-methoxy-diphenylamine-Abstract 2-sulfo-acid (V), methyl ether of 4-amino-4'-Cont'd methoxy-diphenylamine-2-sulfo-acid (VI), 2amino-7-methoxy-phenothiazine-9-dioxide (VII), 4-amino-4'-methoxy-diphenylamine-2-carboxylic acid (VIII), anilide of 4-amino-4'-methoxydiphenylamine-2-carboxylic acid (IX), methyl ether of 4-amino-4-methoxy-diphenylamine-2-carboxylic acid (X), acridone-like compound (XI) and picrate of I (XII). XI is soluble in 2/6 Cara: $\mathbf{E} - 6$

YEAGIUL Cathagan, Analytical Chemistry. General Problems E Abs. John : Not Zear - Kalm., No 9, 1989. No. 15039 Author Institute. Title: Orig Pab. ethanol; II, VI, VII, IX and X - in ethanol and diluted HCl; III, IV, V and VIII - in etha-Abstract Cont'd nol, diluted HCl and alkali; XII - in ethanol, water and diluted HCl. Solutions of leuko compounds of the enumerated ORI are colorless or have a weak yellow color. During the action of oxidizers in an acid solution, ORI first give a blue or violet-blue, and then a red product of oxidation. All derivatives of I are suitable for the indication of oxidation-reduction 3/6 Card:

	E - 7	
Care:	4/6	;
Absyruat. Contid	Processes which take place in an acing the end of titration is determined a change from the colorless form of OHII, IV, V, VI and VII possess a lightion curve with a maximum within 570 During oxidation of II, VIII, IX, X, XII, forms are produced with a viole their maxima of light-absorption are 500-530 mp. At pH 2, I, II, IX and X a stable oxidation-reduction potenti	y the I to blue. ht-absorp610 m \(\text{m} \) XI and t hue, and between
Orig Pub.	:	
Archor Institut. Title	•	5039
	: Per Wir - Row., No. 1, 1959,	۲020
Catagory	Figurials I and witted the istry. General problem.	E

locuntry : EUNGARY Cathering : Analytical Chemistry. General Problems E Abs. Jour : Ref Zhan - Wath., No 5, 1959, No. 15039 Aushor Institut. Titl. Orly Pub. : Abstrent : the color intensity of these ORI does not change in the course of 10 minutes. ORP of Cont'd III, IV, V, VI and VII slowly changes with time; at the same time, a gradual weakening occurs, followed by a disappearance of the color. ORP of VIII, X and XI are very unstable. Taking into account the instability of the ORP of many derivatives of I, the authors consider the ORI of II, III, IV, V, VI, VII, IX and XII to be the most acceptable. The value of SORP Stad: 5/6

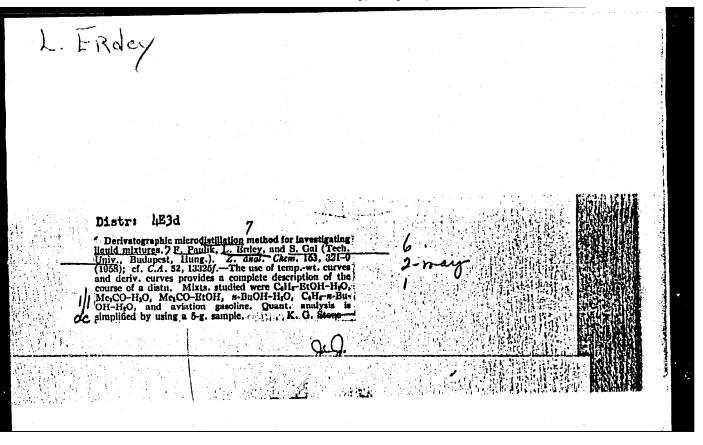
Sauntry : HUNGARY Ostor - y : Analytical Chemistry. General Problems E 65. Jours of extract - 7036., No. 5, 1959. 36, 15039 Auchor Institut. ; Title Original. Absorbet : of the ORI studied depends on the pH. At pH 2, values of SORP for II-X and XII are equal, res-pectively, to 553, 673, 669, 670, 678, 680, 692, 642, 693 and 587 mv. In the opinion of the Cont'd authors, the quoted values of SORP provide the approximate characteristic of the intensity of the attraction or repulsion of the electrons by the corresponding substitutes. The number of electrons which take part in the oxidation-reduction process for ORI, which are derivatives of I, should be equal to 2 .-- N. Polyanskiy Card: 6/6 E - 8

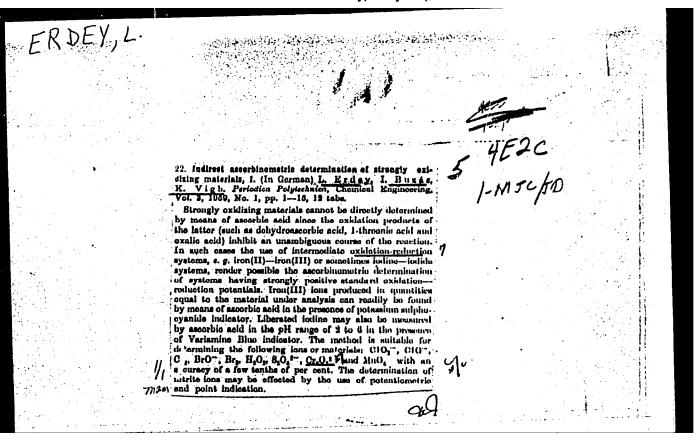


ERDEY, L.				:	
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	Deschiteta exchange	7	and Eva		
	oridizing agent that is ea	reactions. L. Erdey Sudapest, Hung. J. Z. 6m haking a dil, soin. of an ani sol. solid KOx (Ox is a re uily detd.), exchange occur tations are derived for anne will be quant. For C	rs with UX		
	going into soln. Equ whether or not the exch	inge will be quant. For C	NT. AgIO		
	going into soln. Equivalent of the exchion Hg([IO _k]) is best. hy best. Ag:CrO _k or PbC _e Rirors are caused by soly	ange will be quant. For C For SO,, BaCrO, or C O4 is used for S in an OA y, of the ppts, and poor equ K, C	il-, Aglo, la(10,), is le- buffer, ill.		

"APPROVED FOR RELEASE: Thursday, July 27, 2000

CIA-RDP86-00513R00041221





ERDEY, L.; MAZOR, L.; MEISEL.

Data on the microdetermination of the sulfur content in organic compounds. p.404.

MAGYAR KEMIKUSOK LAPJA. (Magyar Kemikusok Egyesulete) Budarest, Hungary. Vol. 1/4, no. 12, Dec. 1959.

Monthly List of East European Accessions. (EEAI) LC Vol. 9, no. 2, Feb. 1960 Uncl.

ERDEY, Lando

Oxidation products of 4-amino-4'-methoxydiphenylamine. Rva Bányai, László Rrdev, and Ferenc Szabadváry (Tech. Univ., Budapest). Acta Chim. Acad. Sci. Hung. 20, 307-20 (1959)(in German).—The polarographic waves and the absorption max. of 4-amino-4'-methoxydiphenylamine (I), of its oxidn. products, and of N-(p-anisyl)-p-benzoquinone dimine (II) proved that II formed in the 2-electronic oxidn. of I. By varying the pH value of the soin., II suffered a change of color, due to the proton affinity of the imino group. The degree of proton affinity of the imino group. The degree of proton affinity of the imino group was established by an optical method and on the basis of the break points of the oxidn.-redn. potential: pH curves. The electrode potential of the oxidn.-redn. system proved to be pH dependent. In a slightly acidic medium, oxidn. took place through a semiquinone intermediate (III), as detd. by using the index potentials. In the oxidn.-redn. potential measurements, the oxidn. agents were: 0.01N Br-H₂O (in acidic soln.) or 0.01N K ferricyanide (in alk. soln.), resp. During the potentiometric oxidn. of I with Br-H₂O at pH 1-6, I gave at first a blue color. By adding Br-H₂O in an amt. corresponding to 2 electrons a violet color arose; and in the presence of strong oxidizing agents (Br-H₂O and Cl-H₂O in great excess), the soln. became red. Over pH 8 the oxidized soln. was continuously yellow.

At pH 1.5-5.5, the 2-electronic oxidn, went through the intermediate III, the stability of which was assured by mesomeric structures. In alk, soin, the oxidn, was direct. At pH 3, a protonated form of II (IV) presented an absorption max, at 580 mµ. The pH region 3-4 was the most favorable for IV (25%). The 2-electronic oxidn, product of I was violet in acidic soin. (absorption max, at 540 mµ), red close to pH 7 (max, at 480 and 540 mµ), and yellow in alk, soin. (max, at 460 mµ); consequently the red color was an mixed one. By polarographic and optical methods, this mixed one. By polarographic and optical methods, this oxidn, product proved to be II. The color change was explained as follows: in aik, soin., II exists as a yellow base; plained as follows: In aik, soln., Il exists as a yellow base; in acidic soln., however, by taking up a proton, Il can exist in the two violet mesomeric forms of IV. The overoxidized product arising from the action of Cl-H₂O contained 3.1% N, no Br; and, probably, it was decompd. Below pH 1, the violet IV became coloriess by decompn. into N-(p-anisyl)-p-benzoquinone imine and NH₂. In weakly acidic medium I took up only one proton, probably on the primary amino group. Over the pH range 1-6, therefore, both the oxidized and the reduced forms of I may exist as univalent actions.

R. Kasztreiner cations.

ERDEY, Laszle, Prof.Dr. (Budapest); PAULIK, Ferenc (Budapest)

Derivategraphic investigation of bauxites; thermic decomposition of hydrargillite. In German. Acta chimica Hung. 21 no.2:205-218 159. (EEAI 9:4)

1. Institute of General Chemistry, Technical University, Budapest.
(Bauxite) (Gibbsite)

12

ERDEY, Laszlo, Prof.Dr. (Budapest XI. Gellert ter 4.); GYIMESI, Jozsef
(Budapest XI. Gellert ter 4.); MEISEL, Tibor (Budapest XI. Gellert ter. 4)

Preparation of some new complex forming compounds and determination of their constants. In German. Acta chimica Hung. 21 no.3:327-332 159. (REAI 9:5)

1. Institute of General Chemistry, Technical University, Budapest. (Complex compounds) (Dissociation)

Distr: hE2c(3)/4E3d

204/60.

Preparation of signs as we complexing agents and the detarmination of their properties. L. Except J. Q. yim sel.
T. Meisell. Mayor Kimich Todgirat, yol. 65, 1050, No.

10, pp. 386—388, 3 figs.

Complexing properties were expected on the basis of practical considerations from the following compounds:
DL-2,3-dihydrexypropylamine, N-dissectic acid, DL-2-arines, N-discetic acid sedium and Legistamio An dissectio sold disedium and the disconsistion domains as awd has the stabilities of the almaliance aritimetal complexes of the analyzed pure materials were determined. It was found that the stability of the alkaliance aritimetal complexes of DL-2,3-dihydraxy, propylamine-N-discett from the amiliary dilesses and includescele acid. The stability of the similarly dilesses and includescele acid. The stability of the complexes of the former compound. The stabilities of the complexes of the former compound. The stabilities of the complexes of the former compound and agreement with the corresponding values of the aritimanionic-N-discetti acid is higher by about one order of magnitude than the gumplex stabilities of the complexes of the prepared new computations of agreement with the corresponding values of the aritimanionic-N-discetti as auxiliary complexing agents, by a withing them in various implements and the stability of the prepared compounds were nonpounds and analysis of the prepared compounds were nonpounds. Conference of the prepared compounds were inferior to those of the prepared compounds were inferior to those of the prepared compounds were inferior to those of the prepared compounds and the compounds is a limited, they can be used only as auxiliary complexing agents.

If the complexing agents are computed as the compounds is a limited, they can be used only as auxiliary complexing agents.

Distr: LE2c

Titrations With Hydrogen Peroxide and Sodium Hypobromite Solutions. L. Erdey
and J. Inczedy (Tech. Univ., Budapest, Hung.). Z. anal. Chem. 166, 110-17 (1959).

-The change from weakly green Ni(OH)2 to black Ni(OH)3 can be used as an indicator
for titrations with 0.1 or 0.01N OBT solns. in weakly basic soln. To det. OBT, add
3 drops 5% NiSO1, soln. (I) and titrate with 0.1N H2O2 to the disappearance of the
black color. To det. S in steel dis-place H2S with HCl in a Schulte app., catch the
H2S in NaOH, oxidize S-- to SO1- with excess OBT, and det. the excess with H2O2.

NH1 salts are detd. by oxidizing NH1+ to N with excess OBT and detg. the excess with
H2O2. OC1- is detd. by adding KBT and titrating with H2O2 and add 3 drops I as indicator. As+++, S--, SO3-, and S2O3- can be titrated directly with OBT soln. in O.1cator. As+++, S--, SO3-, and S2O3- can be titrated directly with OBT soln. in O.1-

(Retyped clipped abstract)
Card 1/1

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ERDEY, Laszlo, r.tag, akademikus

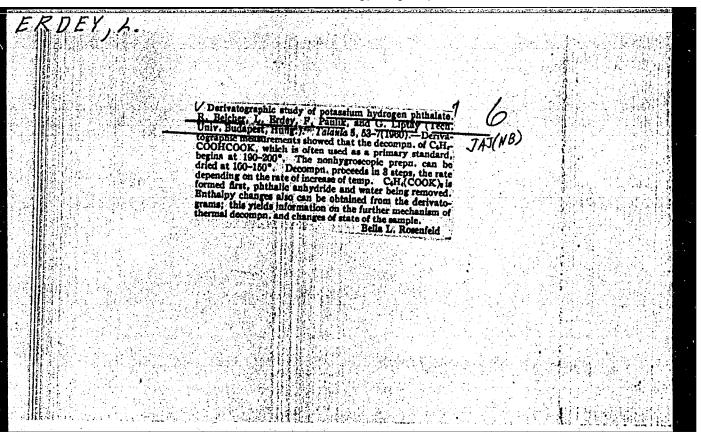
The situation of analytical chemistry and the main trends of its development. Kem tud kozl MTA 14 no.2:213-226 *60. (ERAI 10:2) (Hungary--Chemistry, Analytic) (Hungarian Academy of Sciences)

ERDEY L. prof. (Budapest XI Gellert ter 4); POLOS, L. (Budapest XI Gellert ter 4)

Contributions to the iodometric end point indication. Periodica polytechn chem 4 no.2:157-162 '60. (EEAI 10:4)

1. Institut fur Allgemeine Chemie der Technischen Universitat, Budapest.

(Iodometry) (Potassium iodide)



ERDEY, Laszlo; GYIMESI, Jozsef; MEISEL, Tibor

Synthesis of some new complex-forming compounds and determination of their constants. Hagy kem folyoir 65 no. 10:386-388 0 159.

- 1. Budapesti Muszaki Egyetem Altalanos Kemiai Tanszeke.
- 2. "Magyar Kemiai Folyoirat" szerkeszto bizottsagi tagja.

Gellert ter 4); GAL, S. (Budapest XI., Gellert ter 4); LIPTAY, G. (Budapest XI., Gellert ter 4); PAULIE. (Budapest XI., Gellert ter 4); PAULIE.

Derivategraphic investigation of ammonium phosphate precipitations. Periodica polytechn chem 5 no.3:209-217 161.

1. Lehrstuhl für Allgemeine Chemie, Technische Universität.

ERDEY, L., prof. (Budapest XI., Gellert ter 4); LIPTAY, G. (Budapest XI., Gellert ter 4); GAL, S. (Budapest XI., Gellert ter 4); PAULIK, F. (Budapest XI., Gellert ter 4)

Thermal investigation of iron (III) hydroxy precipitations. Periodica polytechn chem 5 no.4:287-303 '61.

1. Lehrstuhl fur Allgemeine Chemie, Technische Universitat, Budapest. 2. Editorial Board member, "Periodica Polytechnica; Chemical Engineering" (for Erdey).

ERDEY, L., prof.dr. (Budapest XI., Gellert ter 4)

"Progress in nuclear energy - analytical chemistry. Editor M.T. Kelley. Reviewed by Prof., dr. L. Erdey. Periodica polytechn chem 5 no.4:360 '61.

1. Lehrstuhl fur Allgemeine Chemie, Technische Universitat, and Editorial Board member, "Periodica Polytechnica; Chemical Engineering."

ERDEV, Laszlo, prof., dr. (Budapest XI, Gellert ter. 4); GIMESI, Otto (Budapest XI, Gellert ter. 4); RADY, Gyorgy (Budapest XI, Gellert ter. 4)

Determination of elementary sulfur in nonaqueous medium. Acta chimica Hung 28 no.1/3:179-185 '61. (EEAI 10:9)

1. Institut fur Allgemeine Chemie der Technischen Universitat, Budapest.

(Sulfur) (Benzene) (Acetone) (Cyanides)

RADY, Gyorgy (Budapest XI, Gellertter 4); GIMESI, Otto (Budapest XI, Gellertter 4);

ERDEY, Laszlo, prof., dr. (Budapest XI, Gellertter 4)

Determination of the total content of lead and lead oxide in lead chromate. Acta chimica Hung 28 no.1/3:237-242 '61. (EEAI 10:9)

1. Institut fur Allgemeine Chemie der Technischen Universitat, Budapest.

(Lead) (Lead oxides) (Lead chromate)

ERDEY, L. prof. (Budapest, XI., Gellert ter 4); INCZEDY, J. (Budapest, XI., Gellert ter 4)

The role of perhydroxyl ions in the reactions of hydrogen peroxide. Periodica polytechn chem 6 no.4:195-202 162.

1. Department for General Chemistry, Technical University, Budapest.

0/002/62/000/009/001/001 D287/D307

AUTHOR:

Paulik, Ferenc, Paulik, Jeno and Erdey, Laszlo

TITLE:

Derivatography

PERIODICAL:

Chemische technik, no. 9, 1962, 533-537

TEXT: The derivatograph, constructed by the authors, is an automatic recording device for the thermal analysis of solid or liquid samples. Weight changes due to heat and the rate at which these changes proceed and the variations in the enthalpy and the temperature of one sample are recorded simultaneously. The relationship between the chemical composition and the crystalline structure of substances can be determined with a higher degree of accuracy than with hitherto used methods; thermal reactions within the sample can also be elucidated by this method. Derivatograms give results obtained during tests on bauxite samples and during the microdistillation of water. The authors refer briefly to previous investigations on minerals, ores, solid fuels and building materials, on the heat-sensitivity of catalysts and thermal proper-

Derivatography

G/002/62/000/009/001/001 D287/D307

ties of synthetics and state that the method should also give satisfactory results during the analysis of multi-component solvent mixtures, ethereal oils and other valuable organic compounds. A detailed description of the apparatus is included. There are 8 figures.

ASSOCIATION:

Institut für Allgemeine Chemie der Technischen Universität, Budapest (Institute for General Chemistry, Technical University, Budapest)

SUBMITTED:

March 13, 1962

Card 2/2

ERDEY, Laselo; GEGUS, Erno; T. VANDORFFY, Maria

Analysis of natural waters by high-frequency titration. Magy kem lap 17 no.6:277-281 Je '62.

1. Budapesti Mussaki Egyetem Altalanos Kemiai Tanszek,

BECK, Mihaly; BITE, Pal; BRUCKNER, Gyozo; CSENTES, Jozsef; CSUROS, Zoltan; DEAK, Gyula; ERDEY-GRUZ, Tibor; ERDEY, Isszlo; FABIAN, Pal; FINALY, Istvan; FODOR, Gabor; FODORNE CSANYI, Piroska; GYORBIRO, Karoly; INZELT, Istvan; KUCSMAN Arpad; MEUMANN, Erno; PUNGOR, Erno; SCHNEER, Anna; SCHUIEK, Elemer; SZABADVARY, Ferenc

Rules for the Hungarian chemical nomenclature and orthography. Kem tud kozl MTA 17 no.1/4:1-292 162.

1. "A Magyar Tudomanyos Akademia Kemiai Tudomanyok Osztalyanak Kozlemenyei" szerkeszto bizottsagi tagja (for Bruckner, Csuros, Laszlo Erdey, G.Fodor, and Schulek). 2. "A Magyar Tudomanyos Akademia Kemiai Tudomanyok Osztalyanak Kozlemenyei" szerkesztoje (for Erdey-Gruz). 3. "A Magyar Tudomanyos Akademia Kemiai Tudomanyok Osztalyanak Kozlemenyei" technikai szerkesztoje (for Finaly). 4. Muvelodesugyi Miniszterium (for Csentes). 5. Magyar Tudomanyos Akademia Helyesitasi Bizottsage (for Fabian). 6. Nehezipari Miniszterium (for Neumann).

ERDEY-GRUZ, Tibor, akademikus; BRUCKNER, Gyozo, akademikus; LENGYEL, Bela; TELEGDY-KOVATS, Laszlo, a tudomanyok doktora; HARDY, Gyula, kandidatus; GERECS, Arpad, akademikus; FOLDI, Zoltan; WOLKOBER, Zoltan; TUDOS, Ferenc, kandidatus; PURMAN, Jeno; KRAUSZ, Imre, kandidatus; ERDEY, Laszlo, akademikus; SCHAY, Geza, akademikus

An account of the 1961 work of the Section of Chemical Sciences, Hungarian Academy of Sciences. Kem tud kozl 18 no.3:343-394 162.

1. Magyar Tudomanyos Akademia Kemiai Tudomanyok Osztalyanak titkara, es "A Magyar Tudomanyos Akademia Kemiai Tudomanyok Osztalyanak Kozlemenyei" szerkesztoje (for Erdey-Gruz). 2. Akademiai levelezo tag (for Lengyel and Foldi). 3. "A Magyar Tudomanyos Akademia Kemiai Tudomanyok Osztalyanak Kozlemenyei" szerkeszto bizottsagi tagja (for Bruckner, Erdey, Foldi, Gerecs, Hardy, Lengyel, Schay, Tudos).

ERDEY, Laszlo, prof., dr. (Budapest, XI., Gellert ter 4); RADY, Gyorgy, dr. (Budapest, XI., Gellert ter 4); GIMESI, Otto (Budapest, XI., Gellert ter 4)

Analysis of lead-containing silver alloys. Acta chimica Hung 32 no.2:151-157 '62.

1. Institut fur Allgemeine Chemie der Technischen Universitat, Budapest. 2. Mitglied der Redaktion, "Acta Chimica Academiae Scientflarum Hungaricae" (for Erdey).

F. (Budapest, XI., Gellert ter 4); PAULIK, (Budapest, XI., Gellert ter 4); PAULIK, J. (Budapest, XI., Gellert ter 4);

Normalizing the conditions in thermoanalitical experiments by means of a derivatograph. Periodica polytechn chem 7 no. 3: 171-175 163

- Lehrstuhl fur Allgemeine Chemie, Technische Universitat, Budapest.
- 2. Mitglied, Redaktionskollegium, "Periodica Polytechnica-Chemical Engineering." (for Erdey).

ERDEY, L., prof., dr. (Budapest, XI., Gellert ter 4); LIPTAY,
G. (Budapest, XI., Gellert ter 4); PAULIK, F. (Budapest,
XI., Gellert ter 4);

Determination of clacite, magnesite and dolomite in presence of each other by means of a derivatograph. Periodica polytechn chem 7 no. 3: 177-184 163

- Lehrstuhl fur Allgemeine Chemie, Technische Universitat, Budapest.
- 2. Mitglied, Redaktionskollegium, "Periodica Polytechnica-Chemical Engineering". (for Erdey).

ERDEY, L., prof. (Budapest, XI., Gellert ter 4); LIPTAY, G. (Budapest, XI., Gellert ter 4);

Derivate graphic investigation of metal anthranilate precipitate... Periodica polytechn chem 7 no. 3: 185-204 '63

- 1. Lehrstuhl für Allgemeine Chemie, Technische Universität, Budapest.
- 2. Mitglied, Redaktionskollegium, "Periodica Polytechnica-Chemical Engineering" (for Erdey).

ERDEY, L:; prof. (Budapest, XI:, Gellert ter 4); GAL, S. (Budapest, XI., Gellert ter 4)

Thermoanalysis of natural and synthetic cryolite. Periodica polytechn chem 7 no. 3: 205-214 163

- 1. Lehrstuhl für Allgemeine Chemie, Technische Universität, Budapest.
- 2. Mitglied, Redaktionskollegium, "Periodica Polytechnica-Chemical Engineering" (for Erdey).

ERDEY, L., prof. (Budapest, XI., Gellert ter 4); GAL, S. (Budapest, XI., Gellert ter 4); PAULIK F. (Budapest, XI., Gellert ter 4); BAUER, J. (Budapest, XI., Gellert ter 4);

> Derivatographic analysis of calcium oxalate hydrates. Periodica polytechn chem 7 no. 3: 215-22 163

- 1. Lehrstuhl für Allgemeine Chemie, Technische Universität, Budapest (for Erdey, Gal and Paulik).
- Chemische Fabrik Gedeon Richter, Kobanya (for Bayer).
 Mitglied, Redaktionskollegium, "Periodica Polytechnica-Chemical Engineering" (for Erdey).

ERDEY, L., prof. (Budapest, XI., Gellert ter 4); LIPTAY, G. (Budapest, XI., Gellert ter 4)

> Derivatographic study of metal pyridine rhodanide precipitates. Periodica polytechn chem 7 no. 3: 223-236 '63

- Department for General Chemistry, Polytechnical
- University of Budapest.

 2. Editorial board member, "Periodica Polytechnica-Chemical Engineering" (for Erdey).

ERDEY, Laszlo, akademikus

An account of the 3d All-Union Conference on Thermography. Kem tud kozl MTA 19 no.3:355-356 163.

l. Budapesti Muszaki Egyetem Altalanos Kemiai Tanszeke; **a Magyar Tudomanyos Akademia Kemiai Tudomanyok Osztalyanak Kozlemenyei** szerkeszto bizottsagi tagha.

VIGH, Katalin; INCZEDY, Janos; ERDEY, Laszlo

Determination of phosphorus content of steel, crude iron and ferrovanadium by the ion exchange resin column. Magy kem folyoir 69 no.2: 73-75 F 163.

1. Budapesti Muszaki Egyetem Altalanos Kemiai Tanszeke. 2. "Magyar Kemiai Folyoirat" szerkeszto bizottsagi tagja (for Erdey).

ERDEY, Laszlo, KOCSIS, Elemer; TAKACS, Jozsef

Air drying by silica [3]. Epuletgepeszet 12 no.3/4:68-72 Je 163.

1. Budapesti Muszaki Egyetem Altalanos Kemiai Tanszek.

ERDEY, Laszlo, prof., dr. (Budapest, Xl., Gellert ter 4); VANDORFFY, MARIA T. (Mrs), dr. (Budapest, XI., Gellert ter 4)

High-frequency titrations with ascorbic acid as volumetric solution. Acta chimica Hung 35 no.4:381-389 '63.

1. Institut fur Allgemeine Chemie der Technischen Universitat, Budapest. 2. Mitglied, Redaktionskollegium, "Acta Chimica Academiae Hungaricae" (for Erdey).

ERDEY, Laszlo, prof., dr. (Budapest, XI., Gellert ter 4); TESY-VANDORFFY, Maria (Mrs) (Budapest, XI., Gellert ter 4)

High-frequency titrations with ascrobic acid as standard solution. Pt. 2. Acta chimica Hung 37 no.1:17-26 '63.

l. Institut fur Allgemeine Chemie der Technischen Universitat, Budfpest; Mitglied, Redaktionskollegium, "Acta Chimica Academiae Scientiarium Hungaricae" (for Erdey).

ERDEY, L.

"Electrochemical reactions; the electrochemical methods of analysis" by G. Charlot, J. Badoz-Lambling, B. Tremillon. Reviewed L. Erdey. Acta chimica Hung 38 no.2:169-170 '63.

1. Mitglied, Redaktionskollegium, "Acta Chimica Academiae Scientiarum Hungaricae."

GYIMESI, Otto (Budapest, XI., Gellert ter 4); RADY, Gyorgy, dr. (Budapest, XI., Gellert ter 4); ERDEY, Inszlo, Dr. prof. (Budapest, XI, Gellert Teru)

Determination of alkali cyanides and selenium by sulphur volumetric solution in nonaqueous medium. Acta chimica Hung 38 no.4:303-309 '63.

1. Institut für Allgemeine Chemie der Technischen Universität, Budapest.

PAULIK, Ferenc (Budapest, XI., Gellert ter 4); BUZAGH, Eva (Mrs); (Budapest, XI., Gellert ter 4); POLOS, Laszlo(Budapest, XI., Gellert ter 4); ERDEY, Laszlo dr., prof. (Budapest, XI., Gellert ter 4).

Derivatographic analysis of barium sulfate precipitates. Pt.1. Acta chimica Hung 38 no.4:311-323 163.

1. Institut fur Allgemeine Chemie der Technischen Universitat, Budapest.

KASA, Imre (Budapest, XI., Gellert ter 4); ERDEY, Laszlo, prof., dr. (Budapest, XI., Gellert ter 4)

Determination of semicarbazide through oxidation in an alkaline solution. Acta chimica Hung 39 no.1:21-25 '63.

1. Institut fur Allgemeine Chemie der Technischen Universitat, Budapest.

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CSUROS, Zoltar; FETRO, Jozsef; KALPAN, Vince; ERDEY, Lasz. PAULIK, Ference

Changes in the catalytic properties of Raney nickel depending on the conditions of its preparation. Magy kem folyoir 70 nc.8:337-348 Ag 164.

1. Chair of Organic Chemical Technology of the Burbapert Technical University. 2. Editorial board member, "Magyar Kemini Felyoirat", Budapest (for Erdey).

KOROS, Endre; PAULIK, Ferenc; ERDEY, Laszlo; RUFF, Ferenc

Thermal decomposition of some cobalt (II)-pyrazine mixed complexes. Magy kem folyoir 70 no.11:468-474 N '64.

1. Chair of Inorganic and Analytic Chemistry, Lorand Ectvos University, Budapest, Chair of General Chemistry, Budapest Technical University, and Chair of Organic Chemistry, Lorand Fotvos University, Budapest. 2. Editorial board member, "Magyar Kemiai Folyoirat" (for Erdey).

ERDEY, Lazzlo; KANTOR, Tibor; KOCSIS, Elemen; TESYME VANDOPPRY, Maria

Quantitative spectrum analysis of metal layers produced by vacuum evaporation. Magy kem folyoir 70 nc.12:557-559 D 164.

1. Chair of General Chemistry of the Budapest Technical University. 2. Editorial Board Member, "Magyar Kemiai Folyoirat" (for Erdey).